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## Notation

$C_p$ = effective heat capacity, J/kg · K
$C_{p_g}$ = gas heat capacity, J/kg · K
$G(t)$ = cumulative gas produced, $m^3$ gas at STP/$m^2$
$H.V.$ = heating value of gas, J/$m^3$ gas at STP
$k$ = effective thermal conductivity, J/s · m · K
$M$ = gas molecular mass, kg/kg-mol
$P$ = gas pressure, Pa
$P_0$ = pressure at $x = 0$, Pa
$P_D$ = dissociation pressure, Pa
$\dot{Q}(t)$ = cumulative heat input, J/$m^2$
$q(x, t)$ = heat flux, J/s · $m^2$
$R$ = universal gas constant, J/kg-mol · K
$St$ = Stefan number, Eq. 22
$STP$ = standard temperature (273.15K), pressure (101,328 Pa)
$T(x, t)$ = temperature, K
$T_0$ = temperature at $x = 0$, K
$T_D$ = dissociation temperature, K
$T_i$ = initial temperature, K
$t$ = time, s
$v_x$ = superficial gas velocity, m/s
$x$ = axial position, m

### *Greek letters*

$\alpha$ = effective thermal diffusivity, $m^2$/s
$\beta$ = coefficient of thermal expansion of gas, $K^{-1}$
$\Delta H_D$ = hydrate heat of dissociation, J/kg
$\epsilon$ = porosity
$\eta$ = similarity variable ($= x/(4\alpha_{II}t)^{1/2}$)
$\kappa$ = permeability, $m^2$
$\mu$ = gas viscosity, Pa · s
$\xi$ = constant
$\rho$ = effective density, kg/$m^3$
$\rho_g$ = gas density, kg/$m^3$
$\rho_H$ = hydrate density, kg/$m^3$
$\omega$ = mass of gas produced per unit mass of hydrate

### *Subscripts*

I = dissociated zone, I
II = hydrate zone, II
$D$ = dissociation
$g$ = gas
$H$ = hydrate
$i$ = initial condition
0 = boundary condition at $x = 0$

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## Errata

In the paper entitled, "A Nonrandom Factor Model for the Excess Gibbs Energy of Electrolyte Solutions" by A. Haghtalab and J. H. Vera (**34**, May, 1988, p. 803), Eqs. A7 and A8 of the Appendix should read:

$$(\ln \gamma_{I\pm})_{D-H} = \frac{-A|Z_{Ai}Z_{Ci}|I^{1/2}}{1 + BI^{1/2}} \qquad \text{(A7)}$$

$$(\ln \gamma_W)_{D-H} = \frac{2AM_W}{(10B)^3}\left[1 + BI^{1/2} - \frac{1}{1 + BI^{1/2}} - 2\ln(1 + BI^{1/2})\right] + \ln\left(1 + \frac{\nu M_W m}{1{,}000}\right) - \frac{\nu M_W m}{1{,}000} \qquad \text{(A8)}$$

All results were calculated with the correct equations.